# Materials:

* Vinegar
* Baking soda
* Balloon
* Tablespoon
* Funnel
* Plastic bottle
* Scale
* Beaker (optional)
* Safety goggles

# Objectives:

* To verify the law of mass conservation
* To practice making both quantitative and qualitative measurements
* To practice using significant figure rules in calculations

# Overview

In the classic baking soda and vinegar reaction, baking soda (sodium bicarbonate) and vinegar (acetic acid) react to form water, sodium acetate, and carbon dioxide gas. The chemical formula for this reaction is

NaHCO3(s) + CH3COOH(l) → CO2(g) + H2O(l) + Na+(aq) + CH3COO-(aq)

(We’ll learn how to read chemical equations later in the course)

In this experiment, you will perform this reaction and use a balloon to capture the carbon dioxide gas formed. You will measure the mass of the entire setup—including gas—both before and after the reaction.

# Hypothesis:

How do you think the mass of all the ingredients will compare before and after the reaction?
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# Procedure:

1. Measure about 4 tablespoons of baking soda into the plastic bottle.[[1]](#footnote-1) Note that this is a **qualitative** measurement: getting the exact amount is not critical.
2. Measure about 4 tablespoons of vinegar into the balloon.
3. Stretch the neck of the balloon over the mouth of the bottle, being careful not to let the vinegar spill into the bottle.
4. Weigh the balloon and bottle with their contents and record the mass in the Data section. This is a **quantitative** measurement, so make sure to write down all the digits shown on the scale.
5. Lift the balloon to allow the vinegar to pour into the bottle. You may need to hold the balloon at the neck to keep it from coming off. The balloon will inflate as it fills with carbon dioxide gas from the reaction.
6. Once the reaction has settled down, weigh the setup again.

# Data and Calculations:

Mass before reaction: \_\_\_\_\_\_\_\_\_\_\_\_

Mass after reaction: \_\_\_\_\_\_\_\_\_\_\_\_

Calculate the percent change in the mass using this equation:

$$\frac{mass after-mass before}{mass before}×100\%=\frac{ - }{ }×100\%=\frac{ }{ }×100\%= \%$$

Remember to keep track of significant figures using first the addition/subtraction rule, then the multiplication/division rule.

# Conclusions:

Compare the masses from before and after the reaction. Does the data confirm your hypothesis? If not, give a possible explanation why.

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Suggest possible improvements to this lab or possible further work to build on what you’ve learned here.

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1. You could also put the vinegar in the plastic bottle and the baking soda in the balloon. Decide as a group which method you want to use. [↑](#footnote-ref-1)